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ADDITIONS AND CORRECTIONS

Vol. 31, 1953, p. 122, line 7 from bottom. The ratio "87/13" should read "13/87".

Vol. 31, 1953, p. 112. Equation [1] should read $c_2 = wD_2 - \frac{kdD_2}{dx}$.

Vol. 32, p. 136, line 16. "In the limit as $x \longrightarrow 0$" should read "In the limit as $\Delta x \longrightarrow 0$".

Vol. 32, pp. 117–128 and pp. 550–556. One of the authors of the two papers (Paul A. Giguère) has submitted the following note:

THE THIRD LAW ENTROPY AND THE MELTING POINT OF HYDROGEN PEROXIDE

The value 55.76 e.u. published recently (Can. J. Chem. 32: 117–128. 1954) for the calorimetric entropy of hydrogen peroxide in the standard state (i.e. ideal gas at 298.16°K. and 1 atm.) is slightly in error. The correction for ideal gas is in fact negligibly small, 0.003 e.u. instead of 0.10 e.u. as shown in Table III. Therefore the third law entropy in the standard state is *55.66 ± 0.12 e.u.* This correction does *not* affect the entropy of the real gas at 25°C., 67.51 ± 0.1 e.u., nor the values calculated by the statistical methods for other temperatures (in Table IV) for which the uncertainty is greater than the present correction.

Similarly the melting point of hydrogen peroxide should be −0.41°C. instead of −0.43°C. (Can. J. Chem. 32: 550–556. 1954). From the triple point reported in the former paper, 272.74°K., one gets 272.75°K. for the melting point at 1 atm., since the coefficient dt/dp, 0.007° atm.$^{-1}$, is naturally positive.